Introduction to Photolysis
AOSC 433/633 & CHEM 433

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Class Web Site: http://www.atmos.umd.edu/~rjs/class/spr2015

Lecture 10

10 March 2015

Importance of Radicals

- With a few exceptions, the only reactions between molecules that proceed at appreciable rates are those involving at least one radical

- Radicals require significant energy to form: a bond must be broken

- Radical formation is tied to absorption of photons that "photodissociate" a compound, leading to radical formation

Initiation

O_2 + photon $\rightarrow$ O + O

Propagation

$O + O_2 + M \rightarrow O_3 + M$

O_3 + photon $\rightarrow O(^1D) + O_2$

$O(^1D) + H_2O \rightarrow OH + OH$

$OH + O_3 \rightarrow HO_2 + O_2$

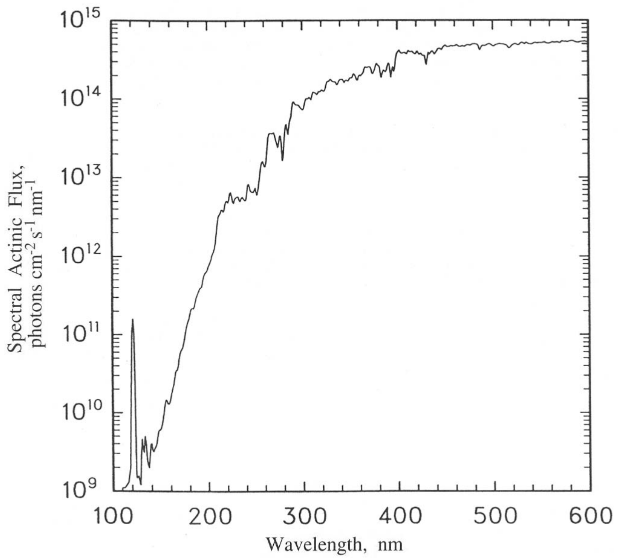
$HO_2 + O \rightarrow OH + O_2$

Termination

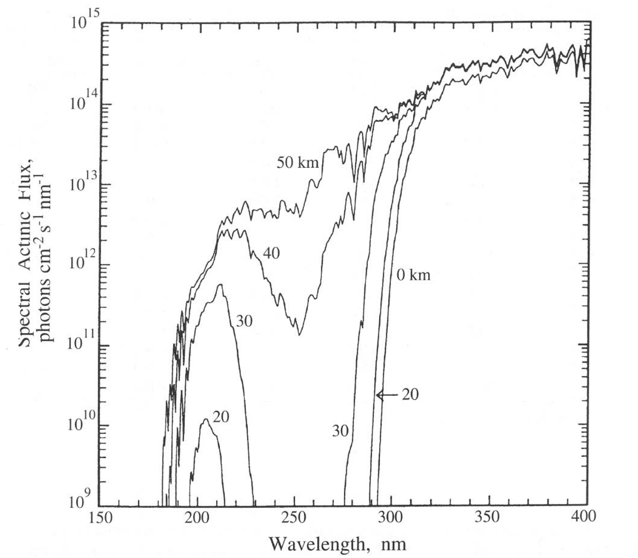
$OH + HO_2 \rightarrow H_2O + O_2$

• Motivation for Today's Lecture:

a) How does atmosphere go from this: **to this ?**



From DeMore et al., *Chemical Kinetics and Photochemical Data for Use in Stratospheric Modeling*, Evaluation No. 11, 1994.



GURE 3.3 Solar spectral actinic flux (photons cm^{-2} s^{-1} nm^{-1}) at various altitudes and at the Earth's surface (DeMore et al., 1994).

From Seinfeld and Pandis, *Atmospheric Chemistry and Physics*, 1998.

b) Biological Effects of UV Radiation

Energetics of Photolysis

$O_3 + h\nu \rightarrow O(^1D) + O_2$

$h\nu$ represents a photon with specific energy.

Let's examine enthalpy of this reaction:

33.9 kcal/mole **104.7 kcal/mole** **22.5 kcal/mole**

$$O_3 \; + \; h\nu \; \rightarrow \; O(^1D) \; + \; O_2(^1\Delta_g)$$ **ΔEnthalpy =**

Photon Energy:

$$\varepsilon = \frac{hc}{\lambda} \quad \Rightarrow \quad \lambda_{\max} = \frac{hc}{\Delta \text{Enthalpy}}$$

For O_3 photo-dissociating to $O(^1D)$:

$$\lambda_{\max} = \frac{hc}{\Delta \text{Enthalpy}} = \frac{2.85 \times 10^4 \text{ kcal/mole nm}}{\Delta \text{Enthalpy}} = \underline{\hspace{3cm}} =$$

Energetics of Photodissociation

$O_3 + h\nu \rightarrow O(^1D) + O_2(^1\Delta_g)$ **ΔEnthalpy = kcal/mole** λ_{max} = **nm**

$O_3 + h\nu \rightarrow O(^3P) + O_2(^3\Sigma_g)$ **ΔEnthalpy = 25.7 kcal/mole** λ_{max} = **nm**

Atomic oxygen: **(Note: you will not be "responsible" for the material below on any exam ☺)**

Ground state – two unpaired electrons in the 2p orbitals: $(1s)^2(2s)^2(2p_1)^2(2p_2)^1(2p_3)^1$

Called ^{3}P:

"**3**" represents 2S+1, where S is spin of all of the unpaired electrons.

There are 2 unpaired electrons, each with spin of ½

Hence, S = 1 and 2S+1 = **3** ⇐ spin angular momentum

P represents orbital angular momentum, found from an electron diagram of filled orbitals:

	$(1s)^2$	$(2s)^2$	$(2p)^4$		
ground state:	↑↓	↑↓	↑↓	↑	↑
			+1	0	-1

$\Lambda = +1 + 1 + 0 - 1 = 1$

Λ	0	1	2	3
Term	S	P	D	F

Excited state – one electron moves from $2p_3$ to $2p_2$: $(1s)^2(2s)^2(2p_1)^2(2p_2)^2$

Called ^{1}D:

"**1**" represents 2S+1, where S is spin of all of the unpaired electrons.

There are no unpaired electrons!

Hence, S = 0 and 2S+1 = **1** ⇐ spin angular momentum

D represents orbital angular momentum, found from an electron diagram of filled orbitals:

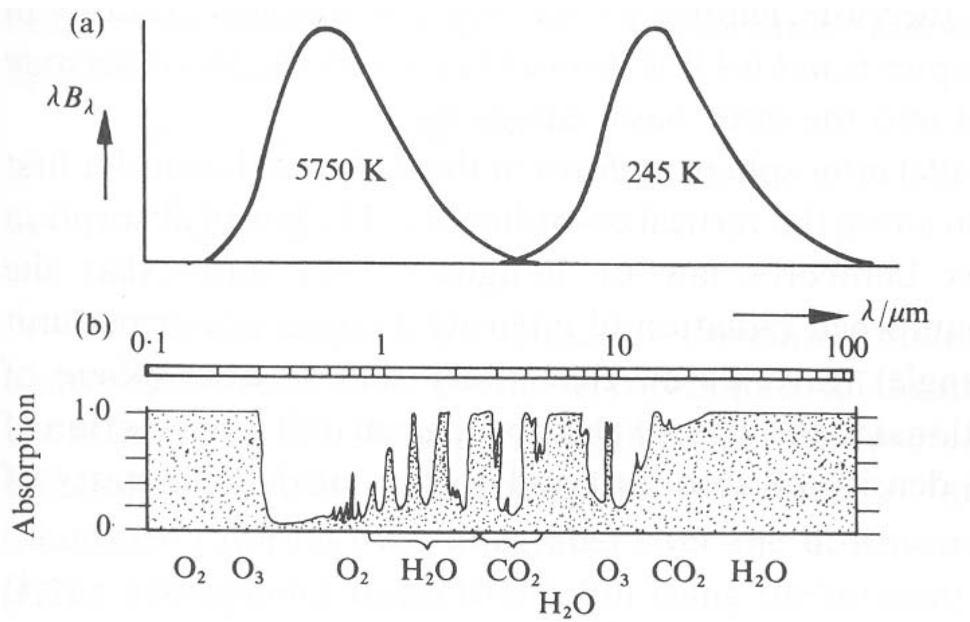
	$(1s)^2$	$(2s)^2$	$(2p)^4$		
excited state	↑↓	↑↓	↑↓	↑↓	
			+1	0	-1

$\Lambda = +1 + 1 + 0 + 0 = 2$

Λ	0	1	2	3
Term	S	P	D	F

Atmospheric Radiation

• Solar irradiance (downwelling) at top of atmosphere occurs at wavelengths between ~200 and 2000 nm (~5750 K "black body" temperature)



Panel (a) : Curves of black-body energy versus wavelength for 5750 K (Sun's approximate temperature) and for 245 K (Earth's mean temperature).

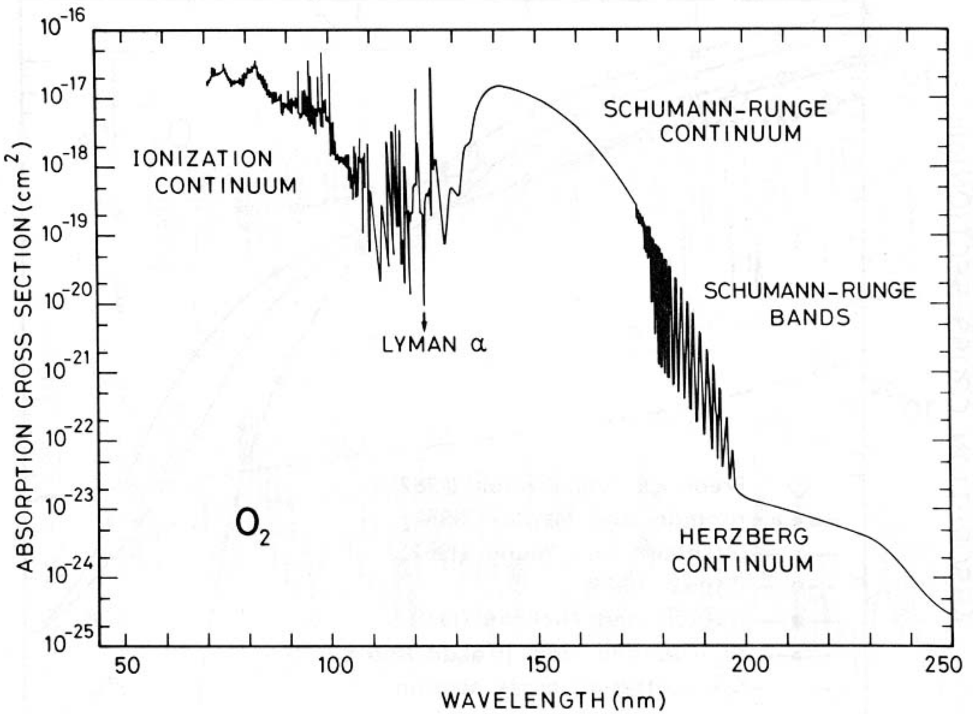
The curves are drawn with equal area since, integrated over the entire Earth at the top of the atmosphere, the solar (downwelling) and terrestrial (upwelling) fluxes must be equal.

Panel (b): absorption by atmospheric gases for a clear vertical column of the atmosphere (1.0 represents complete absorption).

From Houghton, *Physics of Atmospheres*, 1991

• Absorption and photodissociation in the UV occurs due to changes in the electronic state (orbital configuration) of molecules

Absorption Cross Section of O_2



From Brasseur & Solomon, *Aeronomy of the Middle Atmosphere*, 1986

- O_2 can not dissociate longward of ~250 nm
- All absorption shown above is dissociative (e.g., leads to production of two O atoms)
- Structure in the O_2 cross section is related to whether the initial transition involves an unbound electronic state (smooth) or involves a specific vibrational level of an electronic state (banded, due to requirement of specific quanta of energy)

Beer-Lambert Law

$$F(z,\lambda) = F_{TOA}(\lambda)\, e^{-\tau(z,\,\lambda)}$$ (TOA : Top of Atmosphere)

where:

$$\tau(z,\lambda) = m \int_z^{\infty} \sigma_\lambda \,[C]\, dz'$$ (τ: optical depth)

F : solar irradiance (photons/cm^2/sec)
σ_λ : absorption cross section
C : concentration of absorbing gas (molecules/cm^3)
m : ratio of slant path to vertical path, equal to $1/\cos(\theta)$ for $\theta < \sim 75°$
θ : solar zenith angle

Governs basics of radiative transfer in the UV and near IR regions

Photolysis Frequency

For a specific spectral interval, the photolysis frequency (***partial J value***) of a gas is given by the product of its absorption cross section and the solar irradiance:

$$J_{gas}(z,\lambda) = \text{Quantum\_Yield}(\lambda)\ \sigma_{gas}(\lambda,T)\ F(z,\lambda)$$

Units: $s^{-1}\ nm^{-1}$

The total ***photolysis frequency*** (***J value***) is found by integrating $J_{gas}(z,\lambda)$ over all wavelengths for which the gas photodissociates:

$$J_{gas}(z) = \int_{\lambda_{min}}^{\lambda_{max}} J_{gas}(z,\lambda)\, d\lambda$$

Units: s^{-1}

$$\text{Rate of Reaction} = \frac{dO_3}{dt} = J\ [O_3];\quad \text{Units of } J \text{ are } s^{-1}$$

More precisely, calculations of photolysis frequencies consider the "spectral actinic flux", which represents the amount of available photons integrated over all angles, rather than "solar irradiance". These two quantities differ because of scattering of solar radiation by gases and aerosols, and reflection of radiation by clouds and the surface.

Optical Depth of O_2 Absorption

Recall the *Beer-Lambert Law*:

$$F(z,\lambda) = F_{TOA}(\lambda)\, e^{-\tau(z,\,\lambda)}$$ (TOA : Top of Atmosphere)

where:

$$\tau(z,\lambda) = m \int_z^{\infty} \sigma_\lambda\,[C]\,dz'$$ (τ: optical depth)

Also:

$$\int_0^{\infty} [O_2]\,dz' \approx 4\times10^{24} \text{ molecules/cm}^2$$

O_2 Optical Depth for $\theta = 0°$, z = 0 km			
	σ_{max} (cm^2)	τ (0 km)	$e^{-\tau\,(0\text{ km})}$
Schumann-Runge Continuum			
Schumann-Runge Bands			
Herzberg Continuum			

Photolysis Frequency of O_2

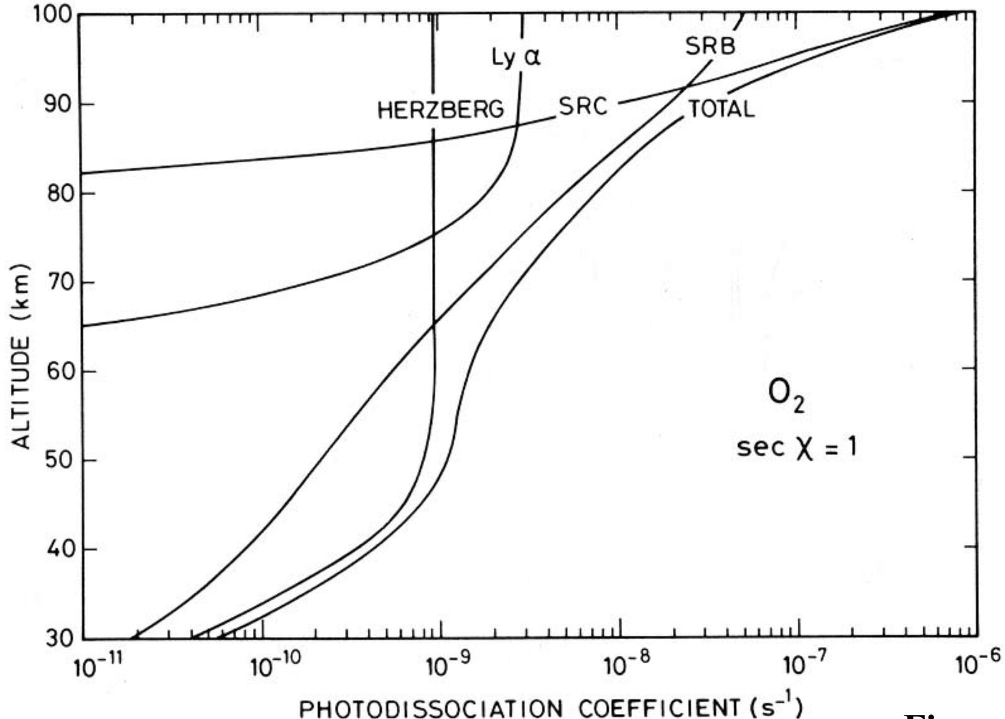


Figure 4.31, Brasseur and Solomon

Where Does Optical Depth = 1.0 for O_2 ?

$$\tau(z, \lambda) = m \int_z^\infty \sigma_\lambda [O_2] \, dz'$$

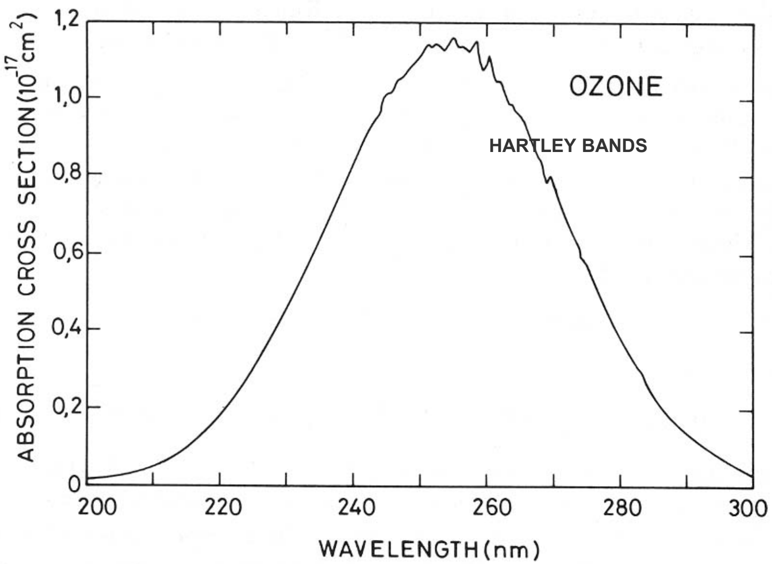
$$\approx \sigma_\lambda \, m \, 4 \times 10^{24} \, e^{-z/H}$$

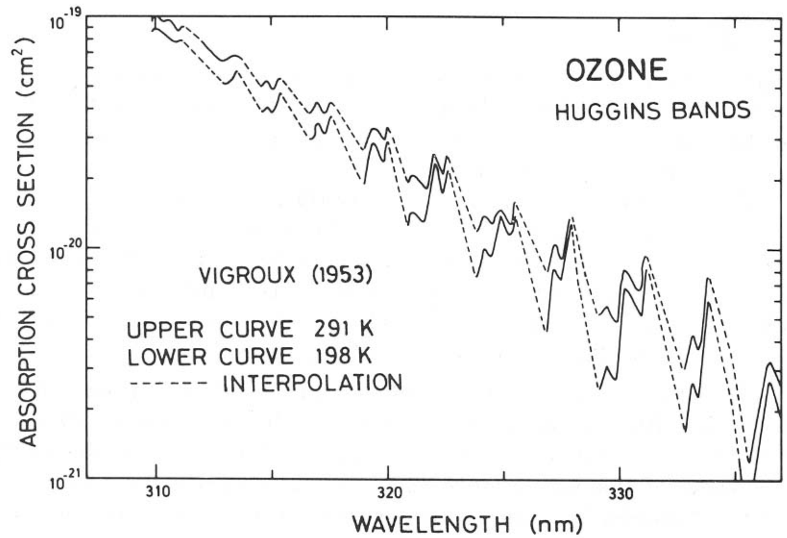
Setting $\tau = 1$ and re-arranging gives:

$$z = H \ \ln (\sigma_\lambda \cdot m \cdot 4 \times 10^{24})$$

Altitude where $\tau = 1$ (for $\theta = 0°$)		
	σ_{max} (cm^2)	z (km)
Schumann-Runge Continuum	10^{-17}	
Schumann-Runge Bands	10^{-20}	
	3×10^{-23}	
Herzberg Continuum	10^{-23}	

Absorption Cross Section of O_3





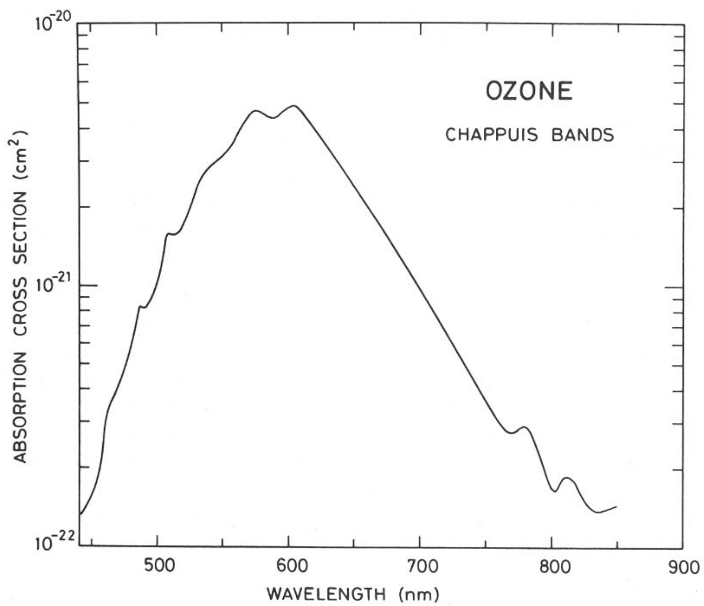


Table 4.6 Theoretical limits corresponding to different photolysis products (nm).

	$O_2(^3\Sigma_g)$	$O_2(^1\Delta_g)$	$O_2(^1\Sigma_g+)$	$O_2(^3\Sigma_u+)$	$O_2(^3\Sigma_u^-)$
$O(^3P)$	1180	590	460	230	170
$O(^1D)$	410	310	260	167	150
$O(^1S)$	234	196	179	129	108

From Brasseur & Solomon, *Aeronomy of the Middle Atmosphere*, 1986

Optical Depth of O_3 Absorption

A typical mid-latitude column abundance for O_3 is 300 Dobson units (DU):

$$1 \text{ DU} = 2.687 \times 10^{16} \text{ molecules/cm}^2; \quad 300 \text{ DU} = 8 \times 10^{18} \text{ molecules/cm}^2$$

Aside: $$\frac{\text{Column O}_3}{\text{Column Air}} = \frac{8\times10^{18}}{2\times10^{25}} = 0.4 \text{ parts per million} \Rightarrow \text{Ozone is a trace species!}$$

O_3 Optical Depth for $\theta = 0°$, z = 0 km				
	σ_{max} (cm^2)	τ (0 km)	$e^{-\tau(0 \text{ km})}$	O_3 Column, $\tau = 1.0$
Hartley (~220 to 280 nm)				
Huggins (~310 to 330 nm)				
Chappuis (~500 to 700 nm)				

Solar Spectral Actinic Flux

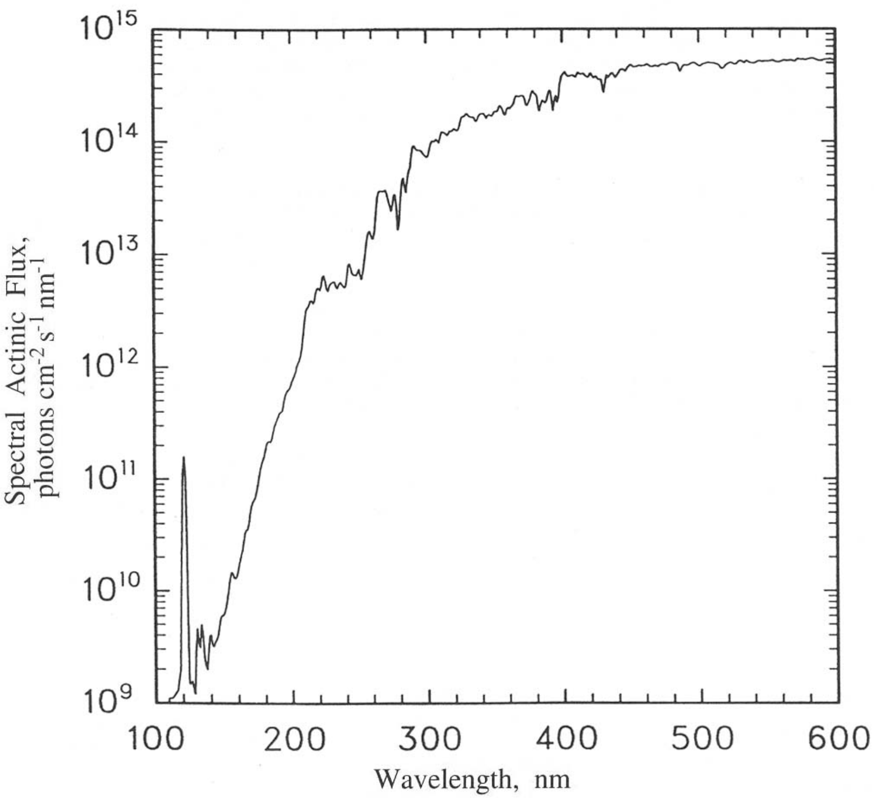
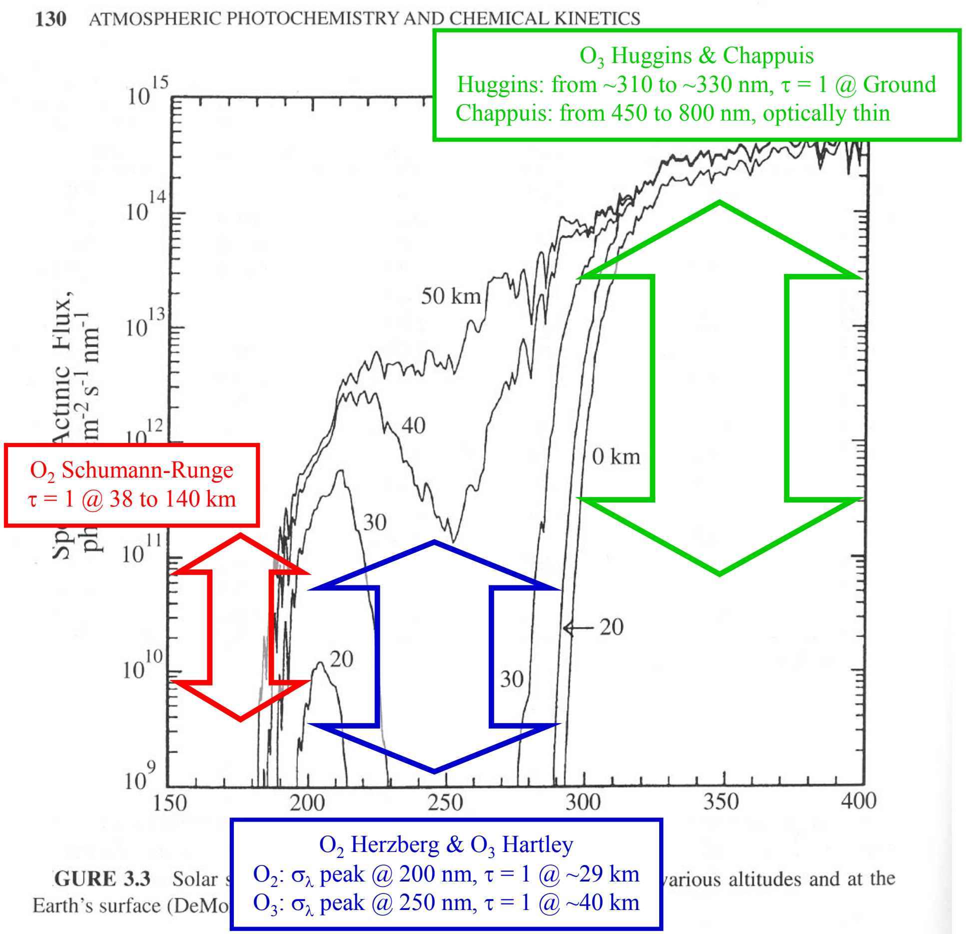


FIGURE 6. Solar spectral actinic flux (photons cm^{-2} s^{-1} nm^{-1}) at the top of Earth's atmosphere.

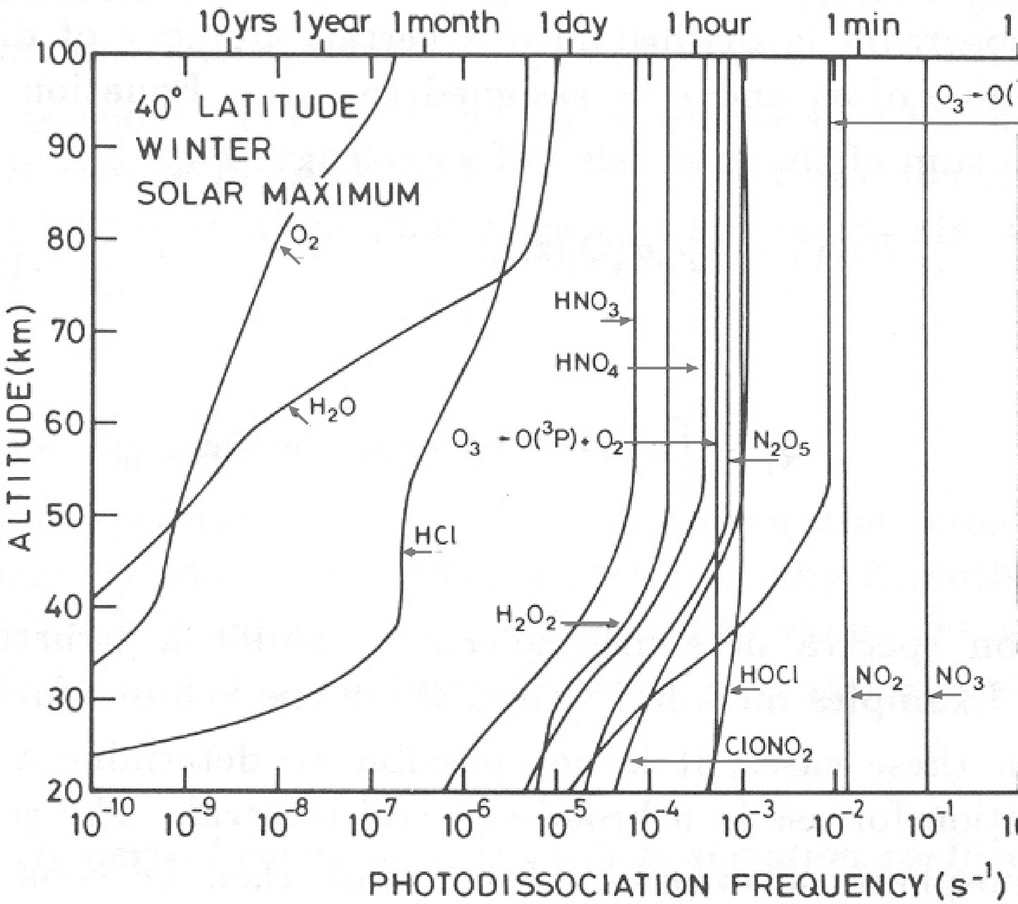


From DeMore et al., *Chemical Kinetics and Photochemical Data for Use in Stratospheric Modeling*, Evaluation No. 11, 1994.

From Seinfeld and Pandis, *Atmospheric Chemistry and Physics*, 1998.

Photodissociation Frequencies

Next goal is to understand:



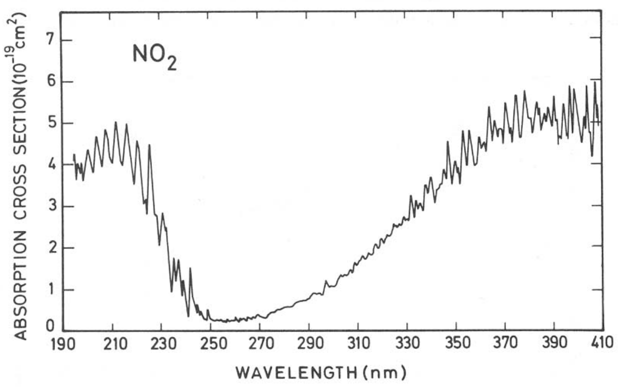
Species	J_∞ (s^{-1})
NO_3	10^{-1}
NO_2	10^{-2}
$O_3 \rightarrow O^1D + O_2$	9.5×10^{-3}
$O_3 \rightarrow O^3P + O_2$	5.9×10^{-4}
N_2O_5	6.6×10^{-4}
HNO_3	7.4×10^{-5}
H_2O_2	2.4×10^{-4}
H_2O	1.0×10^{-5}
$ClONO_2$	10^{-3}
HNO_4	4.5×10^{-4}
$HOCl$	10^{-3}
HCl	3.0×10^{-6}

FIGURE 4.58 Photodissociation frequencies for numerous important atmospheric species.

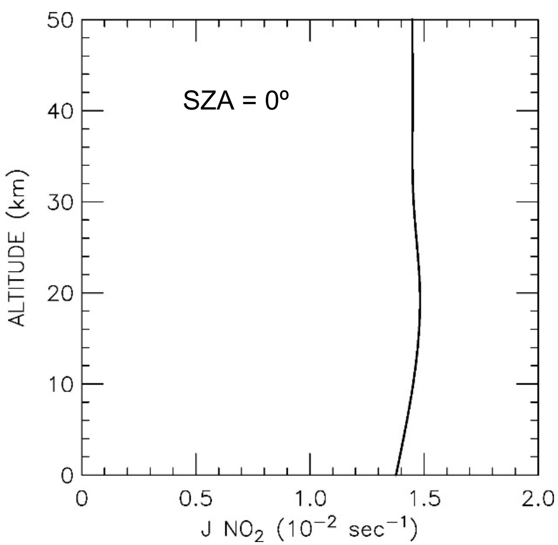
From Brasseur & Solomon, *Aeronomy of the Middle Atmosphere*, 1986

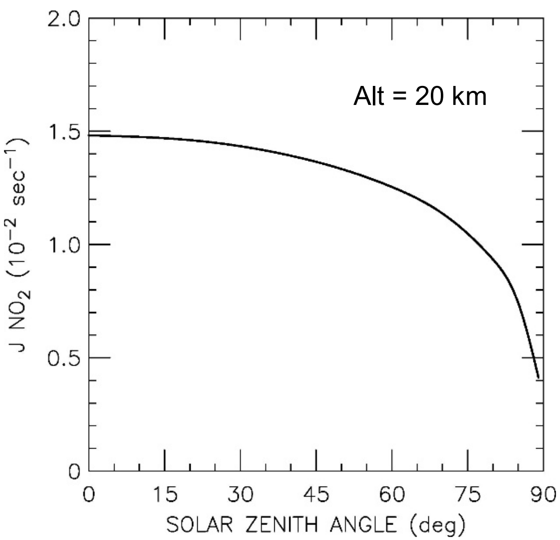
NO_2 Photolysis

The majority of NO_2 photolysis occurs longward of 300 nm, where the atmosphere is optically thin with respect to absorption by O_3 and O_2:

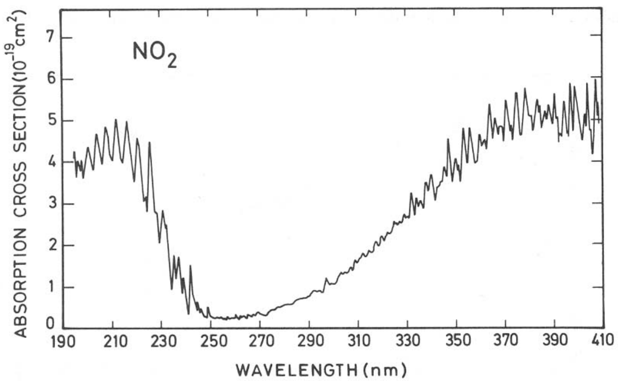


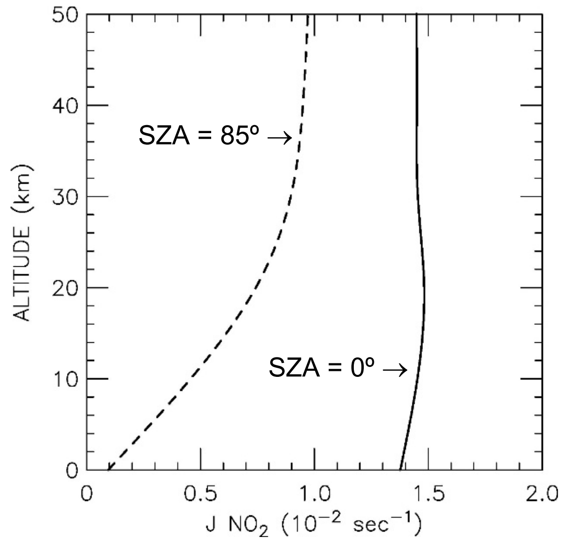
leading to a value for J_{NO2} that is nearly <u>independent</u> of height and SZA:

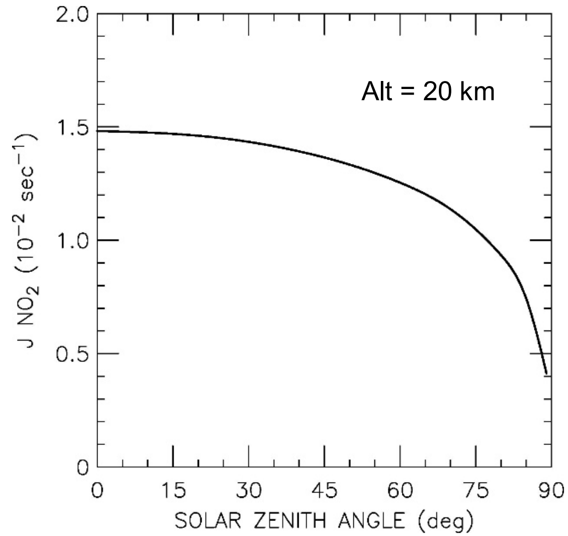


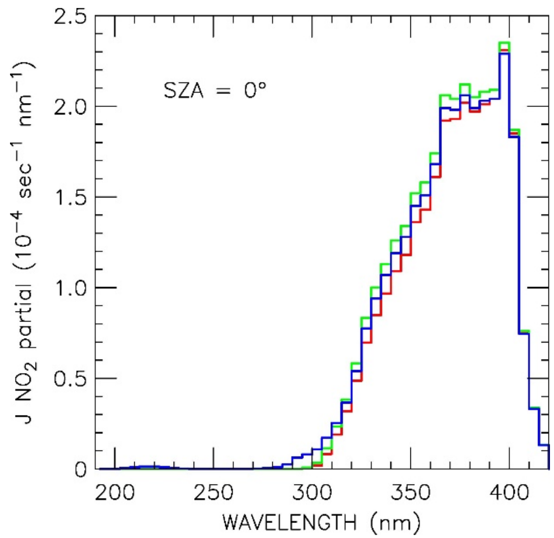


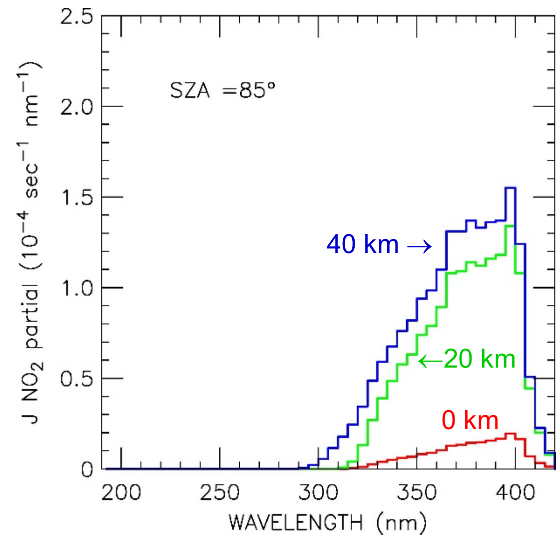
NO_2 Photolysis





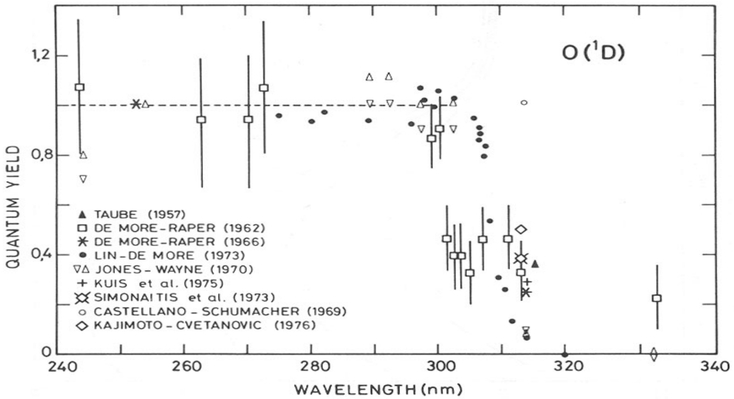


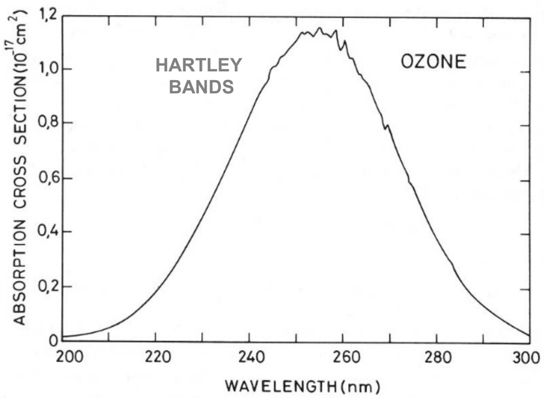


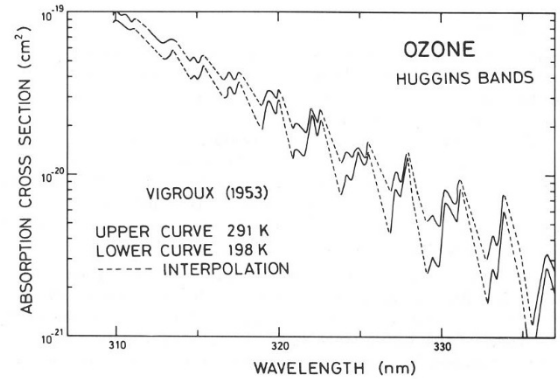


$O_3 \rightarrow O(^1D)$ Photolysis

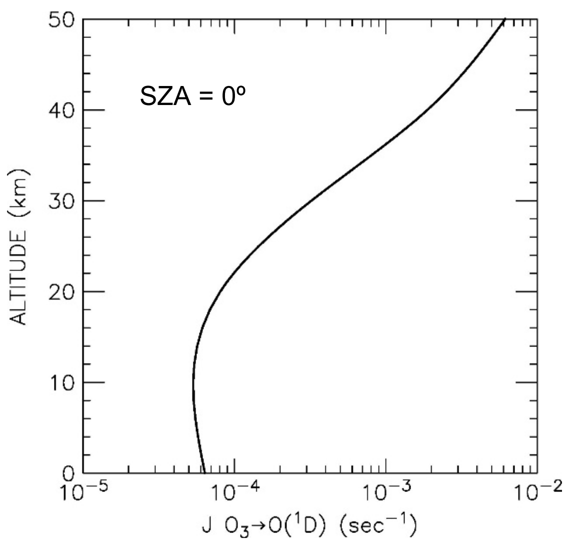
The production of $O(^1D)$ from photolysis of O_3 occurs shortward of 320 nm, where the atmosphere is basically optically thick with respect to absorption by O_3:

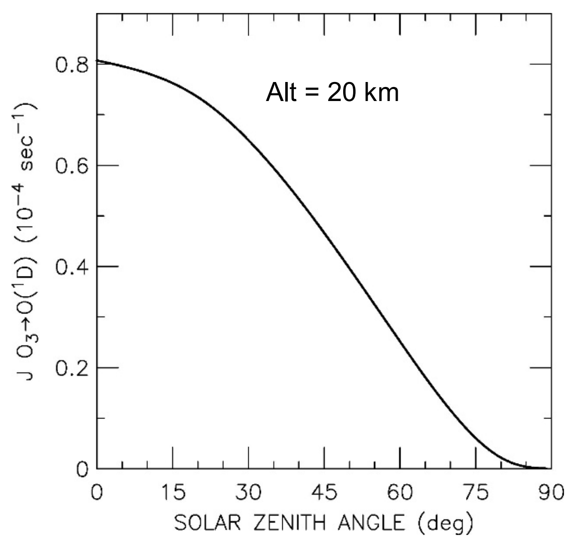




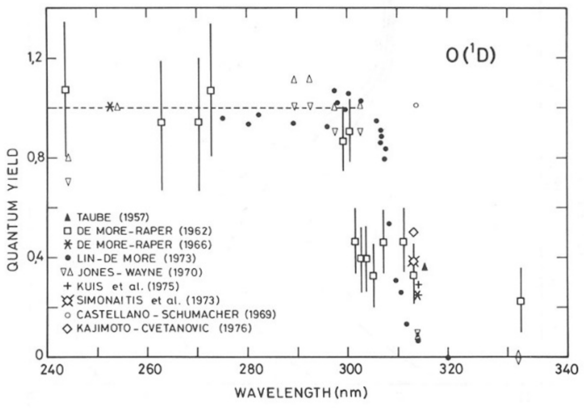


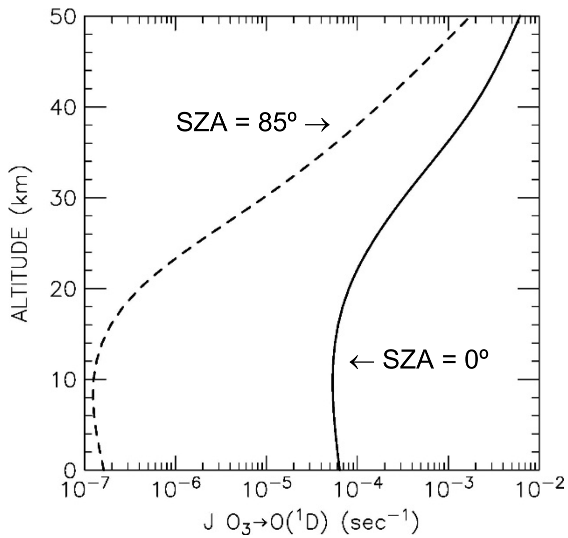
leading to a value for $J_{O3 \rightarrow O(1D)}$ that is <u>dependent</u> on height and SZA:

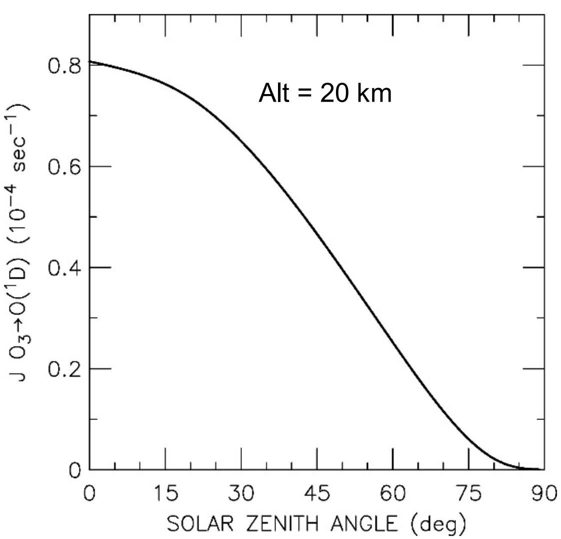


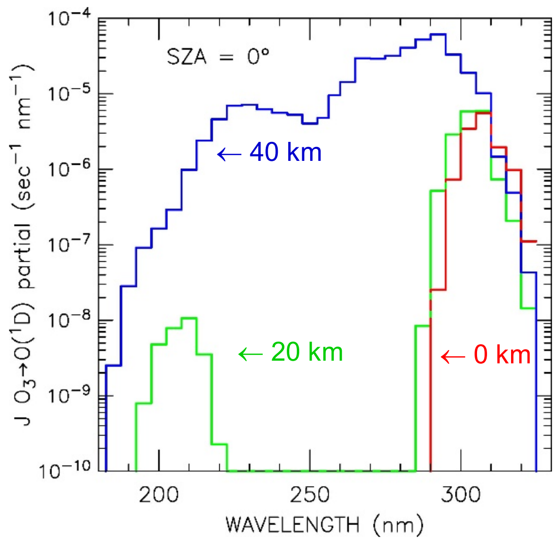


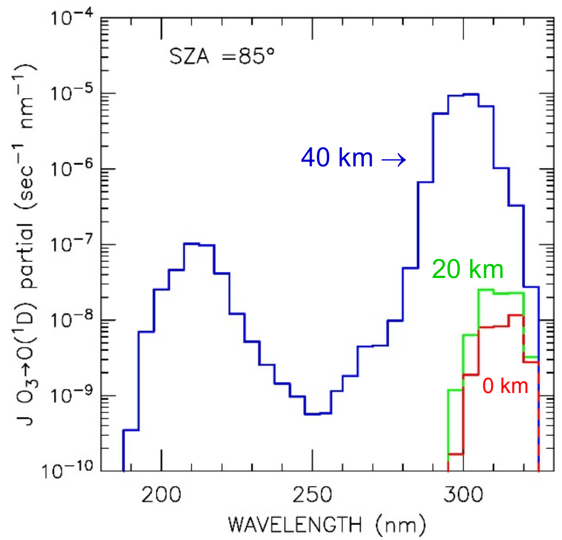
$O_3 \rightarrow O(^1D)$ Photolysis





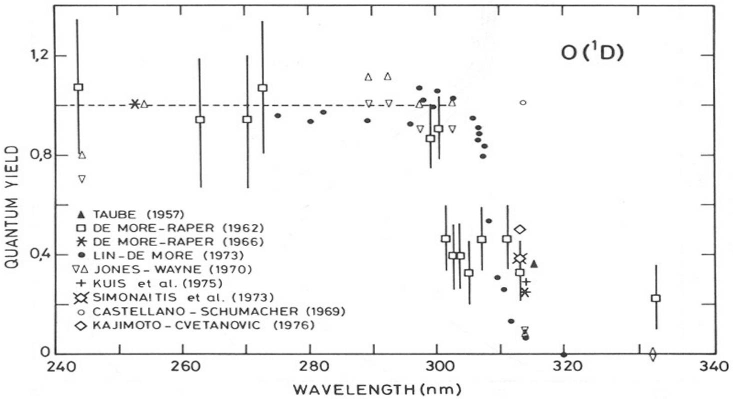


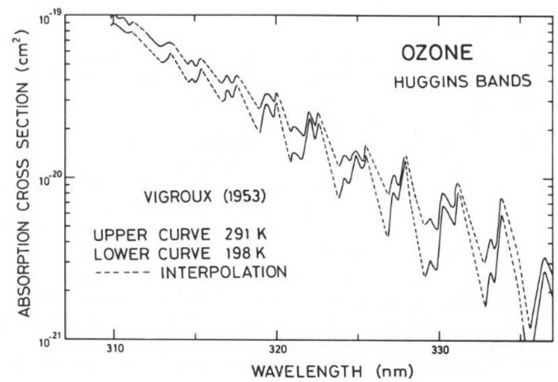


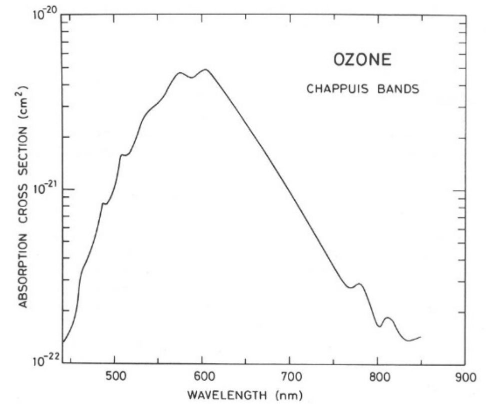


$O_3 \rightarrow O(^3P)$ Photolysis

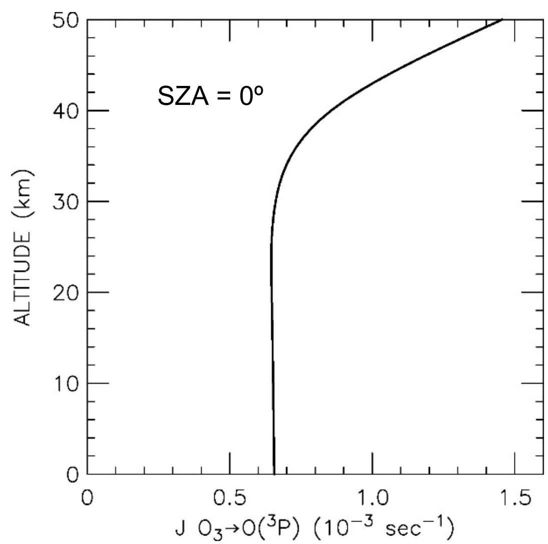
The production of $O(^3P)$ from photolysis of O_3 occurs mainly longward of 500 nm, where the atmosphere is optically thin with respect to absorption by O_3:

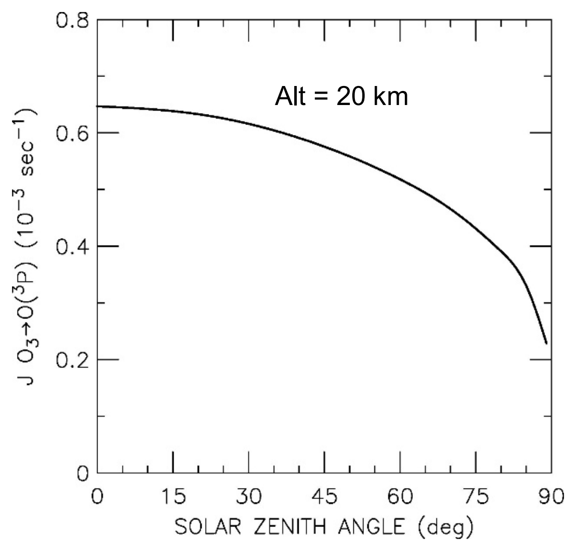




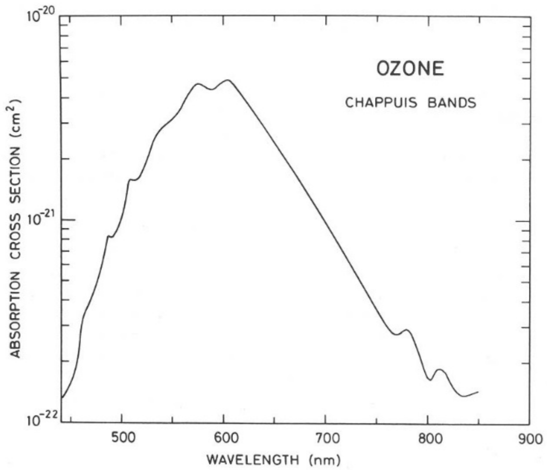


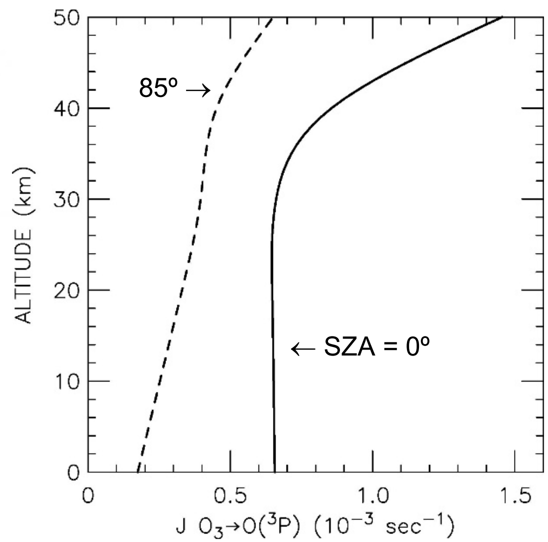
leading to a value for $J_{O3\rightarrow O(3P)}$ that is essentially <u>independent</u> of height and SZA:

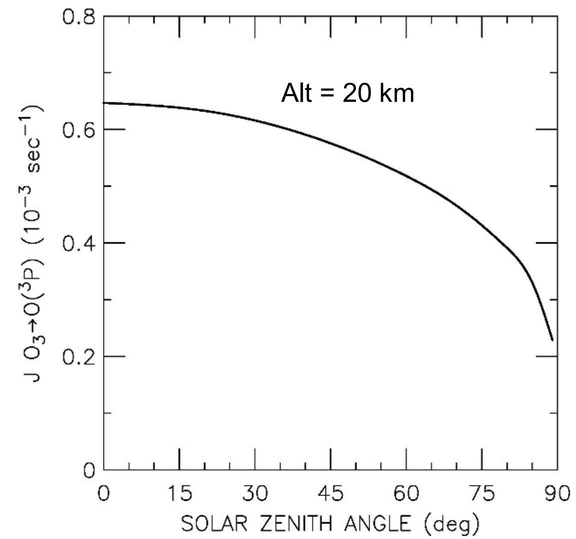


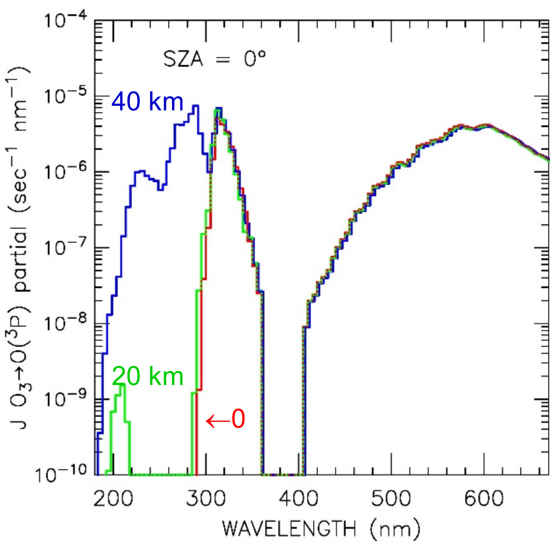


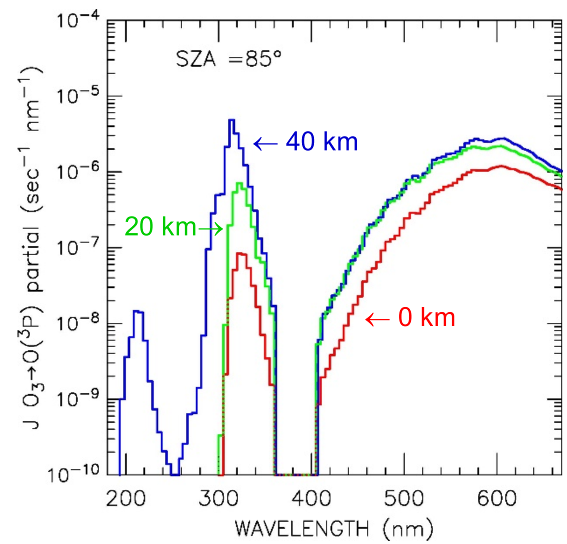
$O_3 \rightarrow O(^3P)$ Photolysis











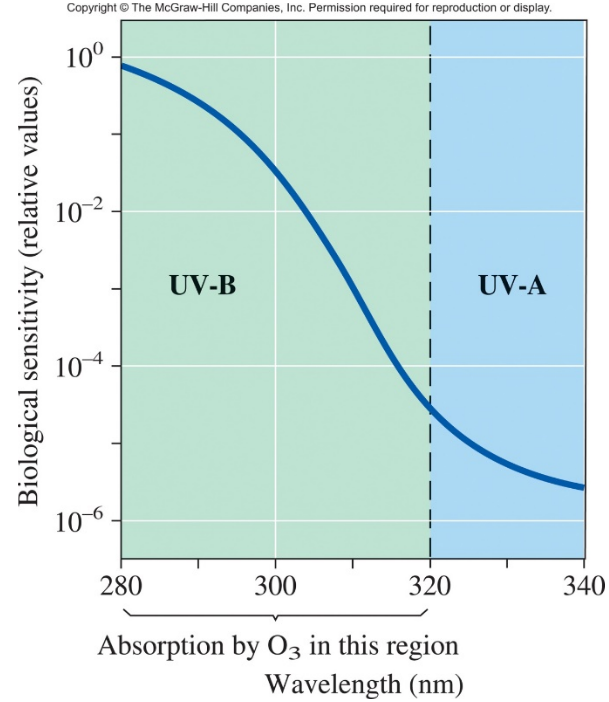
Biological Effects of UV Radiation

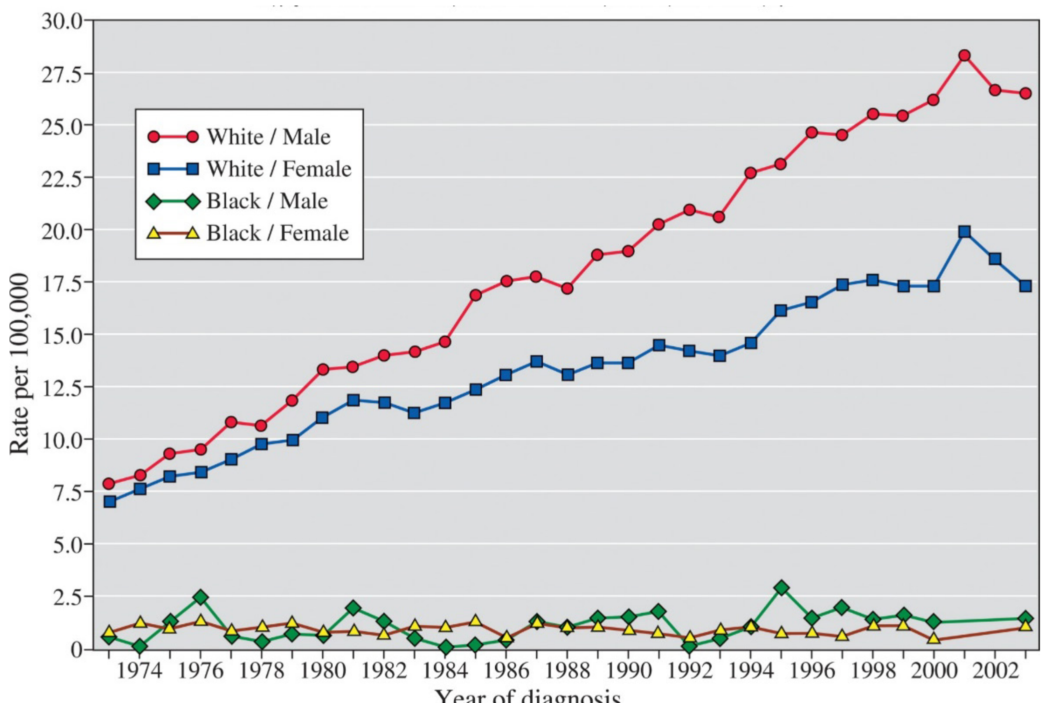
Table 2.4 Types of UV Radiation

Type	Wavelength	Relative Energy	Comments
UV-A	320–400 nm	Lowest energy	Least damaging and reaches the Earth's surface in greatest amount
UV-B	280–320 nm	Higher energy than UV-A but less energetic than UV-C	More damaging than UV-A but less damaging than UV-C. Most UV-B is absorbed by O_3 in the stratosphere
UV-C	200–280 nm	Highest energy	Most damaging but not a problem because it is totally absorbed by O_2 and O_3 in the stratosphere

Chemistry in Context

Figure 2.11, Chemistry in Context





Increase in incidence of melanoma skin cancer, U.S.
Figure 2.12, Chemistry in Context

Biological Effects of UV Radiation

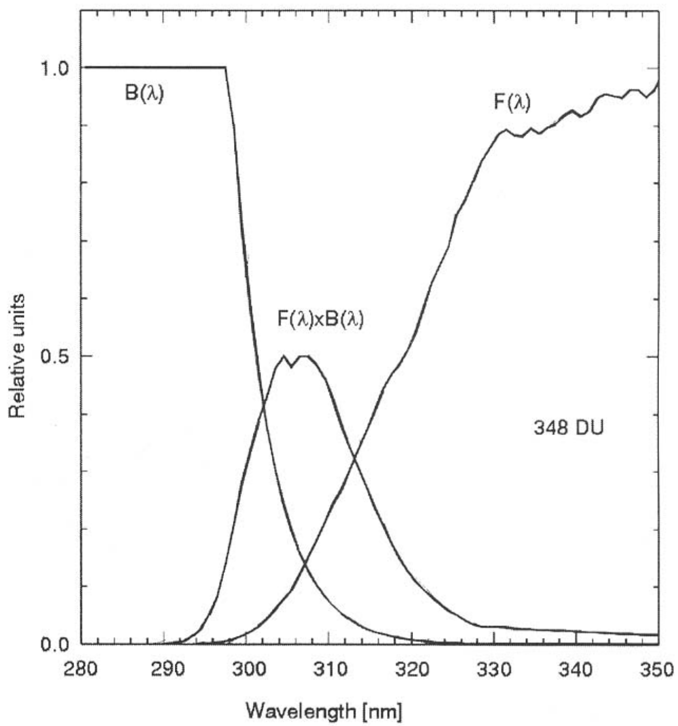


Fig. 1. Biologically active UV radiation. The overlap between the spectral irradiance $F(\lambda)$ and the erythemal action spectrum $B(\lambda)$ given by McKinlay and Diffey [6] shows the spectrum of biologically active radiation, $F(\lambda)B(\lambda)$. The area under the product function $F(\lambda)B(\lambda)$ is the biologically active dose rate. For a total ozone column of 348 DU.

⇐ From Mandronich et al., *J. Photochemistry and Photobiology*, vol. 46, pg. 5, 1998

Humans are:

- strongly affected by exposure to UV-C radiation (100 to 280 nm)
- moderately affected by exposure to UV-B radiation (280 to 315 nm)
- weakly affected by exposure to UV-A radiation (315 to 400 nm)

http://www.who.int/uv/uv_and_health/en

The "biologically active dose rate" maximizes in the UV-B region at ~305 nm, where σ_{O3}= 3 × 10^{-19} cm^2 ⇒ τ (0 km) = 2.4 (for O_3 column= 300 DU)